Intercalation of Lithium Ions from Gaseous Precursors into $\beta$-MnO$_2$ Thin Films Deposited by Atomic Layer Deposition

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ABSTRACT: LiMn$_2$O$_4$ is a promising candidate for a cathode material in lithium-ion batteries because of its ability to intercalate lithium ions reversibly through its three-dimensional manganese oxide network. One of the promising techniques for depositing LiMn$_2$O$_4$ thin-film cathodes is atomic layer deposition (ALD). Because of its unparalleled film thickness control and film conformality, ALD helps to fulfill the industry demands for smaller devices, nanostructured electrodes, and all-solid-state batteries. In this work, the intercalation mechanism of Li$^+$ ions into an ALD-grown $\beta$-MnO$_2$ thin film was studied. Samples were prepared by pulsed LiO$^+$Bu and H$_2$O for different cycle numbers onto about 100 nm thick MnO$_2$ films at 225 °C and characterized with X-ray absorption spectroscopy, X-ray diffraction, X-ray reflectivity, time-of-flight elastic recoil detection analysis, and residual stress measurements. It is proposed that for <100 cycles of LiO$^+$Bu/H$_2$O the Li$^+$ ions penetrate only to the surface region of the $\beta$-MnO$_2$ film, and the samples form a mixture of $\beta$-MnO$_2$ and a lithium-deficient nonstoichiometric spinel phase Li$_x$Mn$_2$O$_4$ ($0 < x < 0.5$). When the lithium concentration exceeds $x \approx 0.5$ in Li$_x$Mn$_2$O$_4$ (corresponding to 100 cycles of LiO$^+$Bu/H$_2$O), the crystalline phase of manganese oxide changes from the tetragonal pyrolusite to the cubic spinel, which enables the Li$^+$ ions to migrate throughout the whole film. Annealing in N$_2$ at 600 °C after the lithium incorporation seemed to convert the films completely to the pure cubic spinel LiMn$_2$O$_4$.

INTRODUCTION

Lithium-ion batteries (LIBs) are used for power storage in portable electronic devices, electric cars, and many other cutting-edge applications. LIBs are rechargeable batteries, which mean that Li$^+$ ions migrate between the anode and cathode reversibly. For the battery to work and maintain its capacity, the electrodes need to be stable when Li$^+$ ions are intercalated to and extracted from the structure upon charge-discharge cycling. With the advent of technologies such as wearable electronics, there is an urgent need for the development of new concepts for the LIBs to meet the increasing demands on energy density, mechanical properties, and capacity retention of the batteries. These demands can be met with nanostructured electrodes and all-solid-state battery architectures.1-4 All-solid-state batteries require the electrode and electrolyte materials to be deposited as thin films. Atomic layer deposition (ALD) is a gas-phase thin-film deposition technique where the substrate is exposed to precursors alternately for self-limiting reaction steps.3 Reactions occur only between the precursor gas molecules and the surface species formed by the other precursor, so that the film thickness increases with an increasing number of reaction cycles. As no chemical reactions take place in the gas phase, the film growth is highly controllable, and the resulting films have excellent uniformity and conformality even on demanding, nanostructured surface topologies.5 These qualities make ALD an ideal technique to deposit electrode materials for LIBs.

The stoichiometric cubic spinel LiMn$_2$O$_4$ crystallizes in a space group Fd$ar{3}$m (Figure 1). The oxygen anions arrange in a cubic closed-packed structure, leaving two possible interstitial sites for the metal cations (denoted by Wyckoff symbols5): tetrahedral 8a, which are occupied by lithium ions, and octahedral 16d, which are filled with manganese cations. LiMn$_2$O$_4$ is a superb cathode material because of its three-dimensional manganese oxide network which enables the intercalation of lithium ions also beyond the 1:1 Li/Mn stoichiometry, resulting in compositions in the range of Li$_x$Mn$_2$O$_4$ ($0 \leq x \leq 2$).9-11 The nonstoichiometric behavior of the oxide improves the specific capacity of the material by...
10%. Furthermore, the low-volume changes upon charge–discharge cycling make it an ideal electrode material for all-solid-state thin-film batteries where the battery structure is rigid.

Generally, the challenge with LiMn$_2$O$_4$ has been that the material experiences capacity fading rather easily because of the distortion of the bonds between oxygen and the Jahn–Teller (JT)-active Mn$^{3+}$ ions. The JT effect deforms the cubic spinel structure to tetragonal or orthorhombic, which can destroy the operation of the LiMn$_2$O$_4$ cathode. Because LiMn$_2$O$_4$ has a critical 1:1 balance of Mn$^{4+}$ and Mn$^{3+}$ ions, the distortion and phase change occur already at relatively low temperatures of 280-293 K when cooling.

However, it was previously reported that the ALD Li$_x$Mn$_2$O$_4$ thin films, prepared by the approach studied also in this paper, show unexpectedly high capacity retention. The capacity of the Li$_x$Mn$_2$O$_4$ electrodes remained high even when they were cycled over a larger voltage range than that is usually applied, which furthermore provides higher gravimetric capacities than what is usually accessible with this material. These properties make the ALD Li$_x$Mn$_2$O$_4$ thin film a very promising cathode material for LIBs and generate questions regarding the mechanism of the intercalation.

Miikkulainen et al. reported a full conversion of MnO$_2$ thin films to Li$_x$Mn$_2$O$_4$ thin films by pulsing LiO$_2$ and H$_2$O sequentially on the MnO$_2$ film at 225 °C. Interestingly, Li$^+$ ions intercalated throughout the whole MnO$_2$ film depth of 100 nm. Normally, the LiO$_2$–H$_2$O process deposits Li$_2$O or LiOH films, but on MnO$_2$ it led to a direct intercalation of Li ions into the MnO$_2$ films. Similar lithiation was attempted also on several other metal oxides ($V_2O_5$, TiO$_2$, Al$_2$O$_3$, ZrO$_2$, Co$_3$O$_4$, Fe$_2$O$_3$, NiO, and MoO$_3$), but was observed to occur only on $V_2O_5$. As ALD thin-film battery materials are likely to be vital contributors to future portable devices, understanding the Li intercalation process into MnO$_2$ is a critical step toward the application of the thin-film LiMn$_2$O$_4$ cathode material. In this work, we aim to understand the intercalation of Li$^+$ ions further by studying the phase and a local crystal structure together with the electronic structure of a series of thin-film samples using X-ray absorption spectroscopy (XAS), X-ray diffraction (XRD), X-ray reflectivity (XRR), and time-of-flight elastic recoil detection analysis (TOF-ERDA) techniques.

## EXPERIMENTAL SECTION

All MnO$_2$ and Li$_x$Mn$_2$O$_4$ samples were deposited in a PICOSUN SUNALE R-150 ALD reactor. The deposition processes have been published earlier by Miikkulainen et al. and Nilsen et al. The depositions took place in two steps: first, MnO$_2$ was deposited using Mn(thd)$_3$ (thd = 2,2,6,6-tetramethyl-3,5-heptanedionato) (Volatec Oy) and ozone (AGA, 99.999%). Lithium was added into the MnO$_2$ films by pulsing LiO$^-$Bu (lithium tert-butoxide) (Acros Organics, 99%) and water alternately. The amount of lithium was controlled by the number of LiO$^-$Bu–H$_2$O cycles. The deposition temperature in both the processes was 225 °C as in the study published earlier. $N_2$ gas (AGA, 99.999%) was used for purging and as a carrier gas.

The MnO$_2$ films were deposited on Si(100) wafers with a pulsing sequence of 0.5 s pulse and 3.0 s purge for Mn(thd)$_3$ and 5.0 s pulse and 5.0 s purge for O$_3$. The pulsing sequence was repeated for 5000 times resulting in around 86–100 nm thick MnO$_2$ film. The source temperature for the Mn(thd)$_3$ powder precursor was 160 °C. O$_3$ was made by feeding O$_2$ (AGA, 99.999%) into a Modular 4 HC ozone generator from Wedeco giving an ozone concentration of about 165 g/m$^3$. A needle valve was used to control the O$_3$ flow into the reactor. A specific nozzle was used in the O$_3$ inlet to ensure gas distribution over the whole wafer.

Lithium was inserted into the film by a pulsing sequence of 0.5 s pulse and 5 s purge for LiO$^-$Bu and 0.1 s pulse and 10 s purge for H$_2$O. The source temperature of LiO$^-$Bu was 160 °C. The LiO$^-$Bu–H$_2$O sequence was repeated for 10, 50, 100, 200, or 300 times, creating a set of samples with increasing lithium concentration. The samples with the highest lithium concentration, that is, the samples with 200 and 300 LiO$^-$Bu–H$_2$O cycles, were also annealed at 400 and/or 600 °C under $N_2$ gas (AGA, 99.999%). For clarity, the samples are named based on their preparation processes (Table 1).

## Table 1. Names and Preparation Methods of the Samples in This Study

<table>
<thead>
<tr>
<th>Name</th>
<th>Preparation Method</th>
<th>Annealing</th>
</tr>
</thead>
<tbody>
<tr>
<td>MnO$_2$</td>
<td>MnO$_2$ film</td>
<td>no</td>
</tr>
<tr>
<td>10Li–MnO$_2$</td>
<td>MnO$_2$ film + 10 LiO$^-$Bu–H$_2$O cycles</td>
<td>no</td>
</tr>
<tr>
<td>50Li–MnO$_2$</td>
<td>MnO$_2$ film + 50 LiO$^-$Bu–H$_2$O cycles</td>
<td>no</td>
</tr>
<tr>
<td>100Li–MnO$_2$</td>
<td>MnO$_2$ film + 100 LiO$^-$Bu–H$_2$O cycles</td>
<td>no</td>
</tr>
<tr>
<td>200Li–MnO$_2$</td>
<td>MnO$_2$ film + 200 LiO$^-$Bu–H$_2$O cycles</td>
<td>no</td>
</tr>
<tr>
<td>300Li–MnO$_2$</td>
<td>MnO$_2$ film + 300 LiO$^-$Bu–H$_2$O cycles</td>
<td>no</td>
</tr>
<tr>
<td>200Li–MnO$_2$–400</td>
<td>MnO$_2$ film + 200 LiO$^-$Bu–H$_2$O cycles at 400 °C for 10 min</td>
<td>no</td>
</tr>
<tr>
<td>200Li–MnO$_2$–600</td>
<td>MnO$_2$ film + 200 LiO$^-$Bu–H$_2$O cycles at 600 °C for 10 min</td>
<td>no</td>
</tr>
</tbody>
</table>

The film thicknesses were analyzed with XRR using a PANalytical X’pert Pro MPD diffractometer or a Rigaku SmartLab diffractometer. Both the diffractometers use Cu Kα radiation with a wavelength $\lambda = 1.5419$ Å. The compositions of the films were characterized with TOF-ERDA. The measurements were done with $^{85}$Br$^+$ ions obtained from a 5 MV tandem accelerator (model EGP-10-II) at the University of Helsinki.

The crystalline phases of the samples were analyzed with XRD. The measurements were done in two geometries. Out-of-plane $\theta$–2$\theta$ diffraction was measured with the PANalytical
diffractometer using programmable divergence and antiscatter slits as focusing optics and a PIXcel 1D detector. In-plane 2θ diatomic was measured using the Rigaku diffractometer and a Rigaku Dtx250 detector with a parallel multilayer mirror and parallel slit collimator optical system. The constant ω and 2θ angles were 0.3°.

The residual stresses of the films were assessed using a Toho Technology FLX-2320-S instrument by measuring the curvature of the sample before and after the film deposition and lithium intercalation. The results were analyzed with a Toho Technology Thin Film Stress Measurement System software.

The electronic structure and the local crystal structure of the samples were studied with the X-ray absorption near-edge structure (XANES) and extended X-ray absorption fine structure (EXAFS) techniques. The measurements were performed for the K- and L-edges of manganese and for the K-edge of oxygen. The K-edge of manganese was studied at the CLÆSS beamline 22 of ALBA synchrotron. During operation, electrons were at 3 GeV in the booster with a current of around 150 mA. The measurements were done at room temperature (27.7 °C) under 10⁻² bar pressure. The primary beam size was 1.5 × 0.5 mm², and it was set to 45° angle with respect to the samples. The samples were measured continuously in total fluorescence yield (TFY) and in total electron yield (TEY) modes. The TFY signal was collected with an XR-100SDD detector located at 90° with respect to the incoming beam and the TEY signal by measuring the drain current of the samples. The data were collected up to k = 18 Å⁻¹ with the scanning sequence: 6400−6520 eV with steps of 1 eV (total of 120 points in 60 s), 6521−6582 eV with steps of 0.2 eV (total of 306 points in 153 s), and 6582.2−7782.8 with steps of 1 eV (total of 1200 points in 600 s). The XANES data analysis was done using the ATHENA program within the IFEFFIT software package together with the curve and peak fitting software Fityk. The EXAFS data were least-square fitted to a theoretical model using ARTEMIS, also within the IFEFFIT package. A k range of 2.912−11.671 Å⁻¹ was used in the Fourier transform (FT) of the EXAFS data. Three powder standards of MnO, MnO₂, and LiMn₂O₄ were used.

The L-edge of manganese and the K-edge of oxygen were studied with XANES at the plane grating monochromator beamline at the laboratory of Physikalisch-Technische Bundesanstalt (PTB) at the synchrotron radiation source BESSY II. The measurement was performed using PTB’s ultrahigh vacuum instrument for X-ray spectrometry and

### Table 2. Results of the Compositional Analysis of the Films

<table>
<thead>
<tr>
<th>Sample</th>
<th>Composition</th>
<th>Li/Mn/O (at. %)</th>
<th>H (at. %)</th>
<th>C (at. %)</th>
<th>F (at. %)</th>
</tr>
</thead>
<tbody>
<tr>
<td>MnO₂</td>
<td>Mn₁.₀₂O₂</td>
<td>0:33:65.6</td>
<td>0.3</td>
<td>0.1</td>
<td>0.2</td>
</tr>
<tr>
<td>10cLi−MnO₂</td>
<td>Li₀.₀₂Mn₁.₀₂O₂</td>
<td>0.4:34:1.63.8</td>
<td>0.5</td>
<td>0.3</td>
<td>0.2</td>
</tr>
<tr>
<td>50cLi−MnO₂</td>
<td>Li₀.₁₀Mn₁.₁₀O₂</td>
<td>1.8:32:1.64.5</td>
<td>0.4</td>
<td>0.1</td>
<td>0.2</td>
</tr>
<tr>
<td>100cLi−MnO₂</td>
<td>Li₀.₁₆Mn₁.₁₆O₂</td>
<td>7.8:31:6.59.7</td>
<td>0.6</td>
<td>0.1</td>
<td>0.2</td>
</tr>
<tr>
<td>200cLi−MnO₂</td>
<td>Li₀.₂₀Mn₁.₂₀O₂</td>
<td>9.7:29:6.59.7</td>
<td>0.3</td>
<td>0.1</td>
<td>0.2</td>
</tr>
<tr>
<td>300cLi−MnO₂</td>
<td>Li₀.₂₅Mn₁.₂₅O₂</td>
<td>14.9:28:4.55.5</td>
<td>0.6</td>
<td>0.2</td>
<td>0.2</td>
</tr>
<tr>
<td><strong>Annealed Samples</strong></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>200cLi−MnO₂−400</td>
<td>Li₀.₃₅Mn₁.₃₅O₂</td>
<td>10.2:30:3.58.4</td>
<td>0.4</td>
<td>0.2</td>
<td>0.2</td>
</tr>
<tr>
<td>200cLi−MnO₂−600</td>
<td>Li₀.₃₅Mn₁.₃₅O₂</td>
<td>11.2:30:1.57.8</td>
<td>0.2</td>
<td>0.2</td>
<td>0.2</td>
</tr>
<tr>
<td>300cLi−MnO₂−600</td>
<td>Li₀.₄₅Mn₁.₄₅O₂</td>
<td>14.1:26:8.58.2</td>
<td>0.4</td>
<td>0.2</td>
<td>0.1</td>
</tr>
</tbody>
</table>

*The composition is an average throughout the whole film thickness.*
related techniques. The three powder standards of MnO2, Mn3O4, and LiMn3O4 that were placed on Si wafer pieces were used. The incident beam angles were 0.5°, 1°, and 45°.

**RESULTS AND DISCUSSION**

The samples were prepared by ALD in two steps. First an MnO2 film was deposited with Mn(thd)3 and O3 as precursors, and second, after characterization of the MnO2 film, lithium was intercalated by pulsing LiO2Bu and H2O alternately so that the composition of the sample changed from MnO2 to LiMn3O4. A set of partially lithiated samples with average stoichiometries through the whole film thickness in the range of Li,MnO2 0 < x < 1 was prepared by increasing the number of LiO2Bu/H2O ALD cycles from 10 to 300.

**Composition, Thickness, and Residual Stress.** Compositional analysis was done with TOF-ERDA. Figure 2 shows the elemental depth profiles, and Table 2 reports the average compositions obtained by integrating over the whole film thickness. The MnO2 film had an Mn/O stoichiometry of 33:66, and the film contained only small amounts of impurities. Fluorine was found from all the samples and it most probably comes from fluoroe lastomer gaskets that were in use in the reactor and in the O3 line. The amount of lithium in the films increased as the number of LiO2Bu/H2O cycles increased, as expected. Stoichiometric LiMn3O4 was achieved after 300 ALD cycles of LiO2Bu and H2O on a 100 nm thick MnO2 film. The depth profiles (Figure 2) reveal that lithium is distributed only in the top part of the films in these samples, the lithium concentrations were 0.6, 3.2, and 9.6 at. %, respectively. In the top parts of the films, the lithium thickness increased by 4–15% during lithium insertion. In the surface roughness of the films varied randomly between 3.3 and 5.5 nm (see Supporting Information Table S1). The depth of the lithium-containing layer increases through the sample series, and in most of the samples the lithiated layer thicknesses are essentially the same as those estimated from the TOF-ERDA depth profiles.

Residual stress was analyzed after the MnO2 film deposition as well as after 300 cycles of lithium insertion. The films for the residual stress measurements were deposited on one side of the 50 mm double-side-polished Si wafers. The residual stress of the MnO2 film was 560 MPa in the tensile direction. The stress decreased compressive and increased to −2300 MPa after the 300 cycles of lithium insertion. As seen in the thickness analysis of 300cLi−MnO2, the film expanded during the lithium insertion, which is thought to be the origin for the high compressive stress.

**Phase Identification, Electronic Structure, and Local Structure.** The phase change from the pyrolusite β-MnO2 to the spinel LiMn3O4 is clearly visible in the XRD and XAS measurements. The samples are divided into three groups with increasing progress in the phase change: when the average composition of the sample is x < 0.5 in Li,MnO2, when x > 0.5 in Li,MnO2, and after the samples are annealed. The following paragraphs explain the details of each of the steps according to the XRD and XAS results.

**Phase Change When the Average Composition is Below x = 0.5 in Li,MnO4.** The in-plane 20 and out-of-plane 0−20 measurements of the samples MnO2−50Li−MnO2 are shown in Figure 3a,b. The samples are mixtures of two phases: the pyrolusite β-MnO2 (ICDD PDF 24-0735), with a space group of P42/mnm, and the spinel Li1−xMn2O4 (ICDD PDF 38-0789), with a space group of Fd3m. In these samples, the reflections from Li1−xMn2O4 are significantly lower in intensity compared to the reflections from β-MnO2. The sharp reflection at 33.0° in the out-of-plane 0−20 measurements is from the Si substrate. In the 0−20 diffractogram of the MnO2 sample, there is a reflection at 41.3°, which is identified as the (200) reflection of the tetragonal pyrolusite MnO2. Additional, the sample has a reflection at 20.3°, which is indexed as (100). The (100) reflection is forbidden for symmetry reasons in the tetragonal pyrolusite space group P42/mnm, but it may appear if the residual stress in the film deforms the structure to be slightly orthorhombic instead of tetragonal. This was suggested earlier by Nilsen et al.29,34 See Supporting Information Figure S2 for more details of the phase identification of the MnO2 sample.

In the beginning of the lithium intercalation, three reflections of Li1−xMn2O4 are somehow visible, the reflection at an angle 44.4° being the strongest one. The reflection is indexed as (400), and it decreases in intensity as the lithium intercalation proceeds. The other two reflections at 17.7° and 35.9° are lower in intensity in these samples. At around 32.6° in the out-of-plane 0−20 measurements of 10cLi−MnO2, there is an additional “impurity” reflection (marked with an asterisk

<table>
<thead>
<tr>
<th>Sample</th>
<th>TOF-ERDA (nm)</th>
<th>XRR (nm)</th>
<th>Film thickness before lithium insertion (nm)</th>
<th>Film thickness after lithium insertion (nm)</th>
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</thead>
<tbody>
<tr>
<td>MnO2</td>
<td>no lithium</td>
<td>no lithium</td>
<td>86</td>
<td></td>
</tr>
<tr>
<td>10cLi−MnO2</td>
<td>~10</td>
<td>~5</td>
<td>100</td>
<td>98 (~2%)</td>
</tr>
<tr>
<td>50cLi−MnO2</td>
<td>~20</td>
<td>~16</td>
<td>96</td>
<td>100 (~4%)</td>
</tr>
<tr>
<td>100cLi−MnO2</td>
<td>~50</td>
<td>~60</td>
<td>97</td>
<td>109 (~11%)</td>
</tr>
<tr>
<td>200cLi−MnO2</td>
<td>whole film</td>
<td>whole film</td>
<td>96</td>
<td>112 (~14%)</td>
</tr>
<tr>
<td>300cLi−MnO2</td>
<td>whole film</td>
<td>whole film</td>
<td>92</td>
<td>108 (~15%)</td>
</tr>
</tbody>
</table>

“The thicknesses of the lithiated layer were analyzed with TOF-ERDA and XRR. The film thickness analysis was done with XRR.”
As shown later, the same reflection is visible in the 100cLi\texttextsubscript{−}MnO\textsubscript{2} sample, and the 200cLi\texttextsubscript{−}MnO\textsubscript{2} sample has a clear additional reflection at 47.4° in the in-plane 2\(\theta\)\texttextsubscript{χ} diffractogram. These reflections are most likely caused by a lithium-rich Li\textsubscript{2−x}MnO\textsubscript{2} (ICDD PDF 38-1282) impurity phase, which has (100) reflection at 32.5° and (012) reflection at 47.6°.

Mn K-edge XANES was measured from every sample in the TEY and TFY (Supporting Information Figure S4) modes. The Mn K-edge EXAFS was measured in the TFY mode. TEY gives information only on the surface of the film, whereas TFY probes the whole film. Although the partially lithiated samples have evident compositional variation in their depth profiles (Figure 2), the TEY and TFY graphs looked surprisingly similar to each other.

The Mn–O, Mn–Mn1, and Mn–Mn2 bonds are readily detected in the Mn K-edge EXAFS of β-MnO\textsubscript{2} and Li\textsubscript{2}MnO\textsubscript{4}. The Mn–O distance means the bond length between Mn and O. The Mn–Mn1 distance is the shortest one between two manganese atoms in the centers of edge-sharing octahedrons, and Mn–Mn2 is the distance between the Mn atoms in the centers of the corner-sharing octahedrons in the β-MnO\textsubscript{2} structure. The longer Mn–Mn2 distance does not exist in the LiMn\textsubscript{2}O\textsubscript{4} structure.

The manganese K-edge XANES and EXAFS measurements support the XRD results, showing that the samples β-MnO\textsubscript{2}–50cLi−MnO\textsubscript{2} have features comparable to the β-MnO\textsubscript{2} reference (Figure 4a,b). Immediate reduction of manganese occurs as soon as lithium is inserted, as the absorption edge shifts to lower energy in the Mn K-edge XANES (Figure 4a). The fine structures of the pre-edge regions of the samples resemble the MnO\textsubscript{2} powder reference (close-up in Figure 4a). The FT EXAFS spectra show that the samples have all the three peaks of Mn–O, Mn–Mn1, and Mn–Mn2 visible in the same way as the MnO\textsubscript{2} reference (Figure 4b).

The L-edge XANES of Mn was studied in addition to the K-edge because the L-edge studies give direct information on the bonding in the samples. The L-edge XANES measures the absorption by the electric dipole-allowed transition from the 2p orbital to the unfilled 3d orbital of a transition metal resulting in an intense and feature-rich L\texttextsubscript{II,III}-edge XANES spectra. In comparison, the K-edge XAS measures the electric
dipole-forbidden transition from the 1s orbital to the 3d orbital, and hence the pre-edge features in the K-edge XANES spectra are usually with lower intensity. The L-edge of Mn was measured at incident angles of 0.5°, 1°, and 45°. The penetration depth of X-rays is dependent on both the photon energy and the incident angle of the X-ray beam. The measurement performed at the 45° incident angle probes the whole film depth, whereas the measurements done with the 0.5° and 1° incident angles are mostly surface-sensitive.

In the Mn L-edge XANES spectrum, the first region comes as a result of Mn^{3+} and the second one as a result of Mn^{4+}.35,36 When the bulk films are compared (Figure 4c), the reduction of Mn is again notable already from the 10cLi–MnO_2 sample onward as the absorption increases at around 641 eV. Similar comparisons of the measurements done with the 1° and 0.5° incident angles are shown in Supporting Information Figure S5a,b.

Phase Change When the Average Composition Exceeds x = 0.5 on Li_xMn_{2-x}O_4. Figure 5a,b show the in-plane 2θ and out-of-plane 2θ–2θ measurements of the samples 100cLi–MnO_2–300cLi–MnO_2. As seen from the TOF-ERDA results, in the samples 200cLi–MnO_2 and 300cLi–MnO_2, lithium has penetrated throughout the whole film thickness. These samples do not anymore have the β-MnO_2 phase in the XRD patterns. The reflections from β-MnO_2 disappear even before the structure has reached the stoichiometric LiMn_2O_4 composition. The two reflections from the lithiated phase shift to lower 2θ values from 17.7° to 17.1° and from 35.9° to 34.7° as the concentration of lithium is increased. In all the samples, the d-spacing of the latter reflection is half of the first reflection, indicating that the reflection on the higher angle is a multiple of the first one. Yet another multiple of the peaks was detected at 75.0° with a d-spacing value of one-quarter of the first peak (see Supporting Information Figure S3). However, these peaks cannot be indexed to the cubic Li_{1−x}Mn_{2-x}O_4 phase. It could be that the samples have a mixture of the cubic and some other lithiated phase, but because no additional phase was recognized in the in-plane diffractograms, even more probable is that the cubic crystal structure is distorted by stress and deformed in one dimension. In addition to stress, the films clearly contain microstrain which is evident from the much
larger full width at half-maximum (fwhm) of the reflection at the higher angle than the fwhm of the one at the lower angle.

Stress and strain could break the $Fd\bar{3}m$ space group symmetry and unveil forbidden reflections. Still more studies are required to reach a complete and thorough phase identification of these samples.

The fine structures of the pre-edge of the Mn K-edge XANES of the samples $100cLi-MnO_2$ and $200cLi-MnO_2$ resemble that of the LiMn$_2$O$_4$ reference (Figure 6a). In addition, the reduction of Mn is again noted as the absorption edge continues its shift to lower energy as the concentration of lithium is increased. The Mn K-edge EXAFS results of all the samples in Figure 6b show only the peaks Mn−O and Mn−Mn1. The absence of the Mn−Mn2 peak supports the XANES results, as it indicates that LiMn$_2$O$_4$ is the dominant phase in these samples. The $300cLi-MnO_2$ sample has an overall different shape in the pre-edge region in the XANES spectra compared to the references and to the earlier published results on the shape of the K-edge for octahedrally coordinated transition metals; therefore, the environment around manganese is probably not octahedral in the $300cLi-MnO_2$ sample.

All the thin-film samples had their absorption edges at lower energies than that of the reference MnO$_2$ powder, and most of the lithiated samples had their absorption edges at lower energies than that of the reference LiMn$_2$O$_4$ powder (Figure 6a). The peaks in the pre-edge region have also higher intensities compared to the references. The higher intensity most likely comes as a result of the JT-active Mn$^{3+}$ ions; so, both features indicate that the average oxidation state of manganese in the samples is lower than +3.5.

The determination of the average oxidation state of manganese will be discussed later in detail, but the presence of the Mn$^{3+}$ ions is noticeable in the EXAFS graphs as well because diminishing of the peaks Mn−O and Mn−Mn1 occurs as a result of the increasing local distortion around the manganese atoms (Figure 6b). Also, the Mn−O bond length seems to increase slightly as the concentration of Li increases in the samples. The lengthening of the bond can occur because of the reduction of Mn$^{4+}$ to Mn$^{3+}$, as the ionic radius of the manganese cation increases and the JT-active Mn$^{3+}$ can distort the structure and thereby lengthen the Mn−O bond even further.

The electrons on the 2p orbitals of oxygen participating in the bonding in the Mn−O octahedra and therefore the oxygen K-edge XANES, which probe the electric dipole-allowed 1s → 2p transition of oxygen, make an interesting addition to the Mn L-edge XAS measurements. The oxygen K-edge was measured for the MnO$_2$, $200cLi-MnO_2$, $300cLi-MnO_2$, $200cLi-MnO_2-600$, and $300cLi-MnO_2-600$ samples at 0.5°, 1°, and 45° incident angles. Both the Mn L-edge and the O K-edge XANES of the $100cLi-MnO_2$−300cLi−MnO$_2$ samples have features that show the presence of Mn$^{3+}$ in the film bulk (Figure 7a,b).
the literature of the Mn L-edge XANES on LiMn$_2$O$_4$ mostly Mn$^{4+}$ was found, but the samples as well as the Mn$_2$O$_3$, Mn$_2$O$_4$, and Li$_2$MnO$_4$ reference powders measured here showed different fine structures compared to the literature results. The literature data were however measured in the TEY mode and are thus surface-sensitive, whereas our data were measured in the T FY mode and are thus bulk-sensitive. The results of the Mn L-edge and O K-edge XANES measured at 0.5° and 1° incident angles are presented in Supporting Information Figures S5 and S6.

All the samples, even MnO$_2$, show Mn$^{3+}$ features on the surface of the film in both the Mn L-edge and in the O K-edge XANES (MnO$_2$ sample in Figure 7c,d and 300cLi–MnO$_2$ in Figure 7ef). These features can be exhibited because of the surface contamination or oxygen vacancies that result in a reduced octahedral symmetry of oxygen atoms at the surface. The 300cLi–MnO$_2$ sample also has an additional resonance at around 532 eV in the O K-edge XANES (Figure 7f). On the basis of the peak shape, the manganese and oxygen atoms do not form an octahedron. The sample may have carbon contamination because according to the literature the K-edge of oxygen bonded to carbon has features similar to that of the MnO references compared to the corresponding unannealed sample. However, because TOF-ERDA shows a low carbon content, the odd fine structure could be a result of oxygen vacancies on the very surface of the film.

The amount of Mn$^{4+}$ and Mn$^{3+}$ in the film bulk was roughly estimated by calculating the different linear combinations of the Mn$^{4+}$ and Mn$^{3+}$ references and comparing them to the measured Mn L$_{III}$-edge XANES (Figure 8). The MnO$_2$ sample resembles the Mn$^{4+}$ reference, as expected, and so do the 10cLi–MnO$_2$ and 50cLi–MnO$_2$ samples. The 100cLi–MnO$_2$ and 200cLi–MnO$_2$ samples match the reference calculated with the Mn$^{4+}$/Mn$^{3+}$ ratio of 80%/20%. The 300cLi–MnO$_2$ sample resembles mostly the reference where the ratio of Mn$^{4+}$/Mn$^{3+}$ is 60%/40%, which is in line with the expectation for this sample with the almost stoichiometric composition of Li$_{1.1}$Mn$_{2.6}$O$_4$.

**Effect of Annealing.** After 300 ALD cycles of LiO’Bu–H$_2$O, the films have a remarkably high compressive stress of 2300 MPa, which is thought to cause a peak shift in the θ–2θ measurements. After annealing at 600 °C for 10 min, the diffraction patterns of 200cLi–MnO$_2$–600 and 300cLi–MnO$_2$–600 match better with the stoichiometric cubic LiMn$_2$O$_4$ reference (ICDD PDF 35-0782) than with the lithium-deficient Li$_{1-x}$MnO$_2$ reference (Figure 9). The in-plane 2θ measurements show how a very clear (111) reflection appears at 18.5° and the already existing peaks shift toward lower angles and thereby match better with the reflections of the reference (Figure 9ac). The two out-of-plane reflections have shifted to higher 2θ angles and can now be indexed as (111) and (311) of the Li$_x$Mn$_{2-x}$O$_4$ phase (Figure 9bd).

In the 200cLi–MnO$_2$–600 sample, the reflections of the MnO$_2$ phase become visible again after annealing (Figure 9b). Similar behavior was not noticed for the 300cLi–MnO$_2$–600 sample. It had, however, an additional reflection at 32.6° (marked with an asterisk in Figure 9d), which is thought to be caused by the earlier discussed impurity phase Li$_{2-x}$MnO$_2$. Annealing at a temperature higher than that used in the lithium intercalation enables the structure to overcome diffusion energy barriers, which in our samples results in a more complete LiMn$_2$O$_4$ crystal structure, but on the other hand also to the recurrence of additional MnO$_2$ and Li$_{2-x}$MnO$_2$ phases.

The Mn K-edge XANES measured in the TEY mode for the annealed samples are shown in Figure 10a,b. (See Supporting Information Figure S4 for the TFY spectra of the annealed samples). The 200cLi–MnO$_2$–400 (not shown) and 200cLi–MnO$_2$–600 (Figure 10a) samples did not have major differences compared to the corresponding unannealed sample. The absorption edge shifted to lower values, but the pre-edge region remained similar in both samples. The Mn K-edge EXAFS spectrum of the 200cLi–MnO$_2$–600 sample resembles the reference Li$_2$MnO$_4$ even better than the corresponding unannealed 200cLi–MnO$_2$ sample (Figure 10c). This result agrees well with XRD in that the annealing completes the phase change.

In the 300cLi–MnO$_2$–600 sample, the e$_g$ resonance has diminished and the pre-edge region resembles more of the MnO reference in the Mn K-edge XANES (Figure 10b). The effect seems to be stronger on the surface of the film (TEY) but it is notable in the bulk measurements (TFY) as well. The sample also has an additional resonance at around 3.5 A in the Mn K-edge EXAFS (Figure 10d), which should not appear for octahedrally coordinated Mn in the spinel Li$_x$Mn$_{2-x}$O$_4$ structure. If the impurity phase was Li$_{2-x}$MnO$_2$ as suggested earlier, the oxidation state of Mn would be close to +2. This would explain the similarity of the pre-edge to the MnO reference in the Mn K-edge XANES (Figure 10e,f). In both cases, the spectra resemble the Li$_2$MnO$_4$ reference.

**Average Oxidation State of Manganese.** The determination of the average oxidation state of manganese was attempted from the XANES results. The energy shifts of several absorption features measured from the reference samples were plotted against the formal oxidation states of manganese. As a result, the half-height of the absorption edge showed the best linear correlation with the oxidation state and was used for determining the average oxidation states in the samples. The analysis was done from both the TEY and TFY measurements, and the oxidation states were similar or at least within the measurement error (see Supporting Information Figure S7 for more details).
The oxidation states show a clear decreasing trend from around +3.7 to around +2.5 as the lithium concentration is increasing (Figure 11). Although the trend is as expected, the measured oxidation states are unrealistically low. However, in addition to the valence of the absorbing atom, structural variations, ligand type, geometry, coordination number, bond lengths, and the covalent or ionic character of the bonds are known to have effects on the chemical shifts in the absorption spectra, which might affect the results and cause the unrealistically low oxidation states.\textsuperscript{46−49} The very low oxidation states of 300cLi\textsubscript{−}MnO\textsubscript{2} and 300cLi−MnO\textsubscript{2}-600 could arise from some unexpected behavior related to 300cLi−MnO\textsubscript{2}, which was also noticed earlier in the shape of the XAS spectrum. More accurate estimation of the oxidation state would require additional studies of the local magnetic properties and site symmetries of manganese, which remains for now beyond the scope of this study.

The determination of the oxidation states was also attempted from the average Mn−O bond lengths obtained from the EXAFS results. The analysis was based on the bond valence model\textsuperscript{50,51} where the oxidation state of an ion is correlated to the nearest neighbor distance in the metal−ligand bond. The oxidation states determined by this method varied in the range of +4 to +3.5, but without any correlation to the lithium concentration.

Proposed Mechanism for the Li\textsuperscript{+} Ion Intercalation. The present results indicate that upon lithiation of the ALD-MnO\textsubscript{2} film, the phase change seems to take place via one spinel lithium manganese oxide phase Li\textsubscript{1−x}Mn\textsubscript{2}O\textsubscript{4} to another one LiMn\textsubscript{2}O\textsubscript{4}. In the beginning of the lithium insertion, that is, in the samples 10cLi−MnO\textsubscript{2} and 50cLi−MnO\textsubscript{2}, lithium penetrated only to the top part of the films where the phase changed to Li\textsubscript{1−x}Mn\textsubscript{2}O\textsubscript{4}. The K absorption edge of Mn shifted to a higher energy as soon as lithium was inserted, which is a result of immediate reduction of Mn\textsuperscript{4+} to Mn\textsuperscript{3+}. The reduction from Mn\textsuperscript{4+} to Mn\textsuperscript{3+} was notable in the fine structure of the L-edge XANES of Mn as well. At this stage, the fine structure of the Mn K-edge XANES still resembled that of β-MnO\textsubscript{2}. The EXAFS of Mn K-edge showed all three bonds Mn−O, Mn−Mn\textsubscript{1}, and Mn−Mn\textsubscript{2} that are characteristic for β-MnO\textsubscript{2}. When the lithium insertion proceeded and the stoichiometry exceeded \(x = 0.5\) in Li\textsubscript{x}Mn\textsubscript{2}O\textsubscript{4}, the out-of-plane orientation of the lithiated phase changed as strong reflections appeared in the \(θ−2θ\) XRD measurements. The TOF-ERDA depth profiles revealed that in this state lithium had diffused throughout the whole film thickness. Features of LiMn\textsubscript{2}O\textsubscript{4} started to appear in the fine structure of Mn K-edge XANES in the sample 100cLi−MnO\textsubscript{2} and onward. The Mn K-edge EXAFS of the samples 100cLi−MnO\textsubscript{2}−300cLi−MnO\textsubscript{2} showed only the bonds between Mn and O and the Mn−Mn\textsubscript{1} distance, which are related to the spinel LiMn\textsubscript{2}O\textsubscript{4}. After annealing at 600 °C, the samples were phase-pure LiMn\textsubscript{2}O\textsubscript{4} according to XRD.

It has been noted in the intercalation studies on powders that as soon as lithium intercalates to β-MnO\textsubscript{2}, manganese reduces from Mn\textsuperscript{4+} to Mn\textsuperscript{3+}.\textsuperscript{37,41,42,52−60} The first intercalation site of lithium is sharing the edges of oxygen octahedron with the Mn\textsuperscript{3+} ions. This intercalation site causes a strong repulsive force and pushes the oxide ions to the cubic close-packed structure, manganese cations to the octahedral 16d interstitial sites, and lithium into the 8a interstitial site, where it is tetrahedrally coordinated. In this structure, the oxygen tetrahedron around the Li\textsuperscript{+} ion is only sharing corners with...
the oxygen octahedron around the Mn$^{3+}$ ions and is hence energetically a more favorable structure.

The proposed mechanism for the intercalation of Li$^+$ ions to $\beta$-MnO$_2$ thin films in this study is as follows. When the lithium concentration is low, that is, when the stoichiometry is $x < 0.5$ in Li$_x$Mn$_2$O$_4$, the lithium ions are distributed only in the surface part of the film, and the phase in this part changes to Li$_{1-x}$Mn$_2$O$_4$. In the rest of the film, the crystal structure resembles still $\beta$-MnO$_2$. When the lithium insertion is continued and $x \approx 0.5$, the phase of the manganese oxide framework appears to change from $\beta$-MnO$_2$ to spinel [Mn$_2$O$_4$]. The phase change occurs even before the composition of the film reaches the stoichiometric spinel LiMn$_2$O$_4$ composition, that is, before 300 LiO$^-$Bu$^-$H$_2$O ALD cycles, presumably because $\beta$-MnO$_2$ can accept only a certain amount of Li$^+$ ions to its square channels, whereas in spinel [Mn$_2$O$_4$], Li$^+$ ions can intercalate via a three-dimensional manganese oxide framework. The phase change enables Li$^+$ ions to intercalate throughout the whole film, but it is not until the samples are annealed that the phase changes completely to phase-pure LiMn$_2$O$_4$. This is because the annealing enables lithium to overcome the diffusion energy barriers and to complete the $\beta$-MnO$_2$/LiMn$_2$O$_4$ phase change.

**CONCLUSIONS**

In this work, the intercalation mechanism of Li$^+$ ions to MnO$_2$ thin films was studied. The films were deposited by ALD in two steps: first, MnO$_2$ films were deposited starting from Mn(thd)$_3$ and O$_3$, and second, lithium was intercalated by pulsing LiO$^-$Bu and H$_2$O on top of the MnO$_2$ films. The phase
change and lithium intercalation were studied by preparing a set of samples with increasing lithium concentration from Li$_{0.025}$Mn$_2$O$_4$ to Li$_{1.1}$Mn$_2$O$_4$. The samples were studied with XAS, XRD, XRR, TOF-ERDA, and residual stress measurements. On the basis of the results, a stepwise model was proposed for the lithium intercalation. Before the lithium concentration reaches $x \approx 0.5$ in Li$_{1-x}$Mn$_2$O$_4$ (<100 LiO$^{2-}$–H$_2$O ALD cycles), the samples are mixtures of $\beta$-MnO$_2$ and Li$_{1-x}$Mn$_2$O$_4$, and after further lithium insertion, the manganese oxide network changes from the tetragonal pyrolusite to the cubic spinel, which enables lithium to penetrate throughout the whole film and the material to achieve the stoichiometric LiMn$_2$O$_4$ structure. Annealing at 600 ºC for 10 min under N$_2$ gas completes the phase change.

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