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1 **When are antiaromatic molecules paramagnetic?**

2

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18

19 **ABSTRACT:** Magnetizabilities and magnetically induced current densities have been cal-
20 culated and analyzed for a series of antiaromatic cyclo[4k]carbons (k=2-11), iso[n]phlorins

1 (n=4-8), expanded porphyrinoids and *meso-meso*, β - β , β - β triple linked porphyrin and iso-
2 phlorin arrays. The cyclo[4k]carbons with k=2-6 are predicted to be closed-shell paramag-
3 netic molecules due to the very strong paratropic ring current combined with its large radius.
4 Larger cyclo[4k]carbons with k=6-11 are diamagnetic, because they sustain a paratropic
5 ring current whose strength is weaker than -20 nA T^{-1} , which seems to be the lower threshold
6 value for closed-shell paramagnetism. This holds not only for cyclo[4k]carbons, but also
7 for other organic molecules like expanded porphyrinoids and oligomers of porphyrinoids.
8 The present study shows that *meso-meso*, β - β , β - β triple linked linear porphyrin and iso-
9 phlorin arrays have a domain-like distribution of alternating diatropic and paratropic ring
10 currents. The strength of their local paratropic ring currents is weaker than -20 nA T^{-1} in
11 each domain. Therefore, linear porphyrin and isophlorin arrays become more diamagnetic
12 with increasing length of the ribbon. For the same reason, square-shaped *meso-meso*, β - β ,
13 β - β triple linked free-base porphyrin and isophlorin tetramers as well as Zn(II) complex of
14 the porphyrin tetramer are diamagnetic. We show that closed-shell molecules with large
15 positive magnetizabilities can be designed by following the principle that a strong para-
16 tropic current ring combined with a large ring-current radius leads to closed-shell paramag-
17 netism.

18

19 **1. INTRODUCTION**

20 The aromaticity concept is manifested by molecular structure,¹ chemical reactivity,² spec-
21 troscopic properties,³ magnetic properties,⁴ and molecular conductivity.⁵ A large number of

1 studies have shown that the degree of aromatic-city can be determined from magnetic
2 shieldings,⁶ ring-current strengths⁷ and other molecular magnetic properties.⁸⁻¹¹ Electron de-
3 localization and pathways of magnetically induced current densities provide information
4 about molecular behaviour in the presence of an external magnetic field (**B**), which is im-
5 portant when designing novel magnetic materials.¹² The magnetic susceptibility (χ) also
6 called magnetizability is a measure of interactions between matter and an external magnetic
7 field (**B**).¹³ The magnetizability (χ) is a second-rank tensor that can be formally divided into
8 diamagnetic (χ^d) and paramagnetic (χ^p) contributions.¹³⁻¹⁷ The magnetically induced current
9 density (**J**), which is the first-order electronic response to an external magnetic field (**B**),
10 can analogously be decomposed into diamagnetic (**J_{dia}**) and paramagnetic (**J_{para}**) current-
11 density contributions.¹⁶

12 Since most closed-shell organic compounds are non-aromatic or aromatic, χ^p is in absolute
13 value often smaller than χ^d . The magnetizability of organic molecules is then diamagnetic
14 corresponding to negative magnetizability values, by definition. The more negative mag-
15 netizability, the stronger diamagnetic is the molecule. When χ^p dominates, the molecule is
16 paramagnetic with a positive magnetizability. Paramagnetic molecules are in general open-
17 shell molecules, whereas closed-shell paramagnetism (CSP) is rare.^{13,18} Diatomic BH and
18 its isoelectronic analogues CH⁺ and BeH⁻ are known to be closed-shell paramagnetic mol-
19 ecules,¹⁹⁻²¹ whereas only a few examples of closed-shell paramagnetic polyatomic mole-
20 cules have been reported.^{22,23}

1 We have shown in a number of recent studies that antiaromatic molecules sustaining a
2 strong paratropic ring current can be closed-shell paramagnetic molecules,²⁴⁻²⁶ which are
3 spectroscopically characterized by intense low-lying magnetic-dipole-allowed electronic
4 transitions. Strong paratropic ring currents also affect the magnetizability, because the an-
5 gular momentum of the paratropic ring current increases the χ^p contribution. Anderson *et.*
6 *al.*²³ recently detected CSP for closed-shell multicharged antiaromatic porphyrin nanorings,
7 which experimentally confirmed that CSP can occur in strongly antiaromatic molecular
8 rings.

9 The article is outlined as follows. In the next section, we briefly discuss the theory of
10 closed-shell paramagnetism. In section 3, the employed computational methods are pre-
11 sented. These methods are applied in section 4.1 to a closed-shell paramagnetic tetra-cat-
12 ionic hexaporphyrin nanowheel, whose magnetizability has been measured experimen-
13 tally. In section 4.2, we study the magnetizability and ring-current strengths of a series of
14 antiaromatic cyclo[4k]carbon rings. In section 4.3 we discuss magnetic properties of por-
15 phyrinoid arrays, while the data for expanded isophlorins and expanded porphyrinoids are
16 presented in sections 4.4 and 4.5, respectively. The aromaticity of the square-shaped
17 [36]octaphyrin(1.1.1.1.1.1.1.1) and cyclo-octafuran(1.1.1.1.1.1.1.1) is discussed in section
18 4.6. The obtained results are summarized and concluded in section 5.

19

20

1 2. THEORY

2 Molecular magnetizability is generally calculated as the second derivative of the electronic
3 energy with respect to the strength of the external magnetic field using the response formal-
4 ism.²⁷⁻²⁹ However, the mechanism of closed-shell paramagnetism can be easier understood
5 from the sum over states expression for the second derivative, where the magnetizability is
6 split into diamagnetic (χ^d) and paramagnetic (χ^p) contributions³⁰

$$\begin{aligned} 7 \quad \chi_{\alpha\beta} = & -\frac{1}{4} \sum_i \langle 0 | r_{Oi}^2 \delta_{\alpha\beta} - r_{Oia} r_{Oib} | 0 \rangle \\ 8 \quad & + \frac{1}{2} \sum_{\substack{i \\ K>0}} \frac{\langle 0 | (\mathbf{r}_{Oi} \times \mathbf{p}_i)_\alpha | K \rangle \langle K | (\mathbf{r}_{Oi} \times \mathbf{p}_i)_\beta | 0 \rangle}{E_K - E_0} \end{aligned} \quad (1)$$

9 The summation in Eq (1) runs over all electrons i and excited states K . $\mathbf{r}_{Oi} = \mathbf{r}_i - \mathbf{r}_O$ is
10 the distance from the common gauge origin, \mathbf{p}_i is the momentum operator, and E_0 and E_K
11 are the energy of the ground and excited states, respectively. The diamagnetic term is always
12 negative and depends only on the ground state, whereas the paramagnetic part contains the
13 orbital angular momentum operator and is always positive. Low-lying excited states with
14 large magnetic transition moments from the ground state have large paramagnetic contri-
15 butions to the magnetizability. These excited states have also been found to be the reason
16 for the antiaromatic character of isophlorin and other antiaromatic porphyrinoids.²⁶
17 Magnetizabilities can also be obtained by integrating the magnetically induced current den-
18 sity multiplied with the vector potential of the external magnetic field:^{16, 31}

$$19 \quad \chi_{\alpha\delta} = -\frac{1}{2c} \varepsilon_{\alpha\beta\gamma} \int r_\beta J_\gamma^{B\delta}(\mathbf{r}) \, d\mathbf{r}, \quad (2)$$

1 where $\varepsilon_{\alpha\beta\gamma}$ is the Levi-Civita tensor, r_β is the vector potential and $J_\gamma^{B\delta}(\mathbf{r})$ is the magneti-
2 cally induced current density susceptibility that we call current density in the rest of this
3 article. When the external magnetic field is applied in z direction, χ_{zz} is obtained as:

$$4 \quad \chi_{zz} = -\frac{1}{2c} \int (x J_y^{Bz}(\mathbf{r}) - y J_x^{Bz}(\mathbf{r})) d\mathbf{r} \quad (3)$$

5 When a planar aromatic molecule lies in the xy plane, $J_y^{Bz}(\mathbf{r})$ flows in the positive y direc-
6 tion for positive x values and in the negative y direction when x is negative, leading to a
7 negative contribution to χ_{zz} . The second term in Eq. (3) is also negative for aromatic mol-
8 ecules. The paratropic ring current of antiaromatic molecules flows in the opposite direc-
9 tion and results in positive contributions to χ_{zz} . The two contributions in Eq. (3) are large
10 for strong ring currents whose radius is large. For ring-shaped molecules the χ_{xx} and χ_{yy}
11 elements of the magnetizability tensor are negative as for most molecules and they scale
12 nearly linearly with the size of the molecule as originally observed by Pascal.³² The average
13 molecular magnetizability is calculated as the trace of the magnetizability tensor:

$$14 \quad \chi = \frac{1}{3} (\chi_{xx} + \chi_{yy} + \chi_{zz}) \quad (4)$$

15 implying that the paramagnetic χ_{zz} contribution of antiaromatic molecules in the xy plane
16 is cancelled by the diamagnetic contributions from $\chi_{xx} + \chi_{yy}$. The magnetizability is pos-
17 itive when χ_{zz} is larger than the absolute value of $\chi_{yy} + \chi_{xx}$. We have previously shown
18 that porphyrinoids become closed-shell paramagnetic when the strength of the paratropic
19 ring current exceeds -20 nA T^{-1} ,²⁴ which holds for molecules with about the same radius of
20 the ring current as isophlorin.

1 3. COMPUTATIONAL METHODS

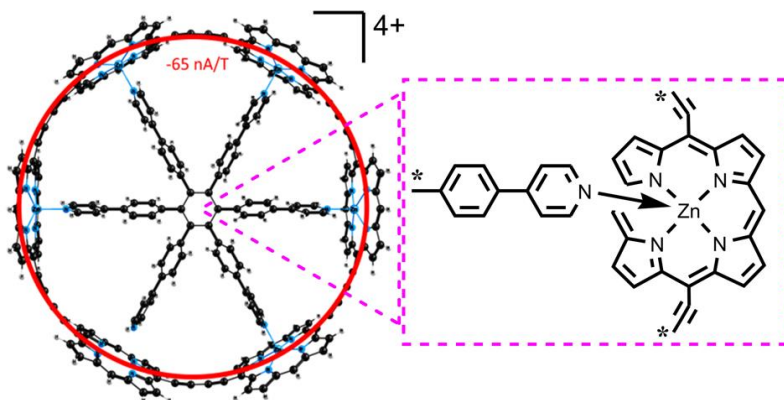
2 The molecular structures were optimized at the density functional theory (DFT) level us-
3 ing the M06-2X exchange–correlation functional³³ and the def2-TZVP basis set³⁴. It is
4 known that this hybrid functional well describe magnetic properties and magnetizabili-
5 ties.³⁵⁻³⁶ The NMR shielding constants were calculated by using the Gauge Including
6 Atomic Orbital (GIAO) method³⁷ at the M06-2X level of theory. These calculations were
7 performed for isolated molecules using the Gaussian 16 program package.³⁸

8 The GIMIC program^{39,40} was used for calculation of the magnetically induced current
9 densities and the integrated current densities. The magnetically induced current densities
10 were calculated using density matrices obtained in the NMR shielding calculations and
11 basis-set data. The magnetic field was oriented perpendicular to the molecular plane. The
12 strengths of the current densities were obtained by numerical integration of the current
13 density passing a rectangular surface perpendicular to the molecule plane. The magnetiz-
14 ability was also calculated using GIMIC.

15 We performed state-specific complete active space self-consistent field (SS-CASSCF)
16 calculations for free-base porphyrin arrays (**H₂P**)_n and isophlorin arrays (**IPh**)_n with
17 n=1,2,3,4. The SS-CASSCF calculations included 8 electrons in 8 molecular orbitals
18 (MO). The weight of closed shell determinant varies in the range of 0.9-0.93. For other
19 determinants, they are less than 0.15 implying that the studied molecules are single-refer-
20 ence systems. The SS-CASSCF calculations were carried out with the def2-TZVP basis
21 set using the Firefly software.⁴¹

1 4. RESULTS AND DISCUSSION

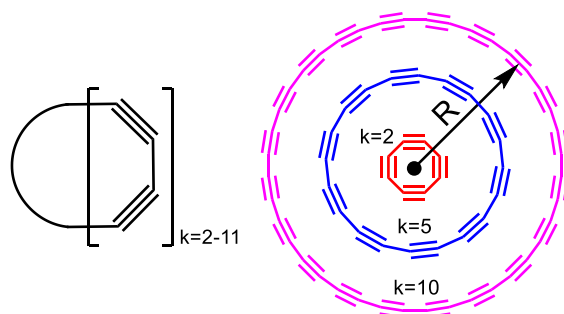
2 **4.1 Hexaporphyrin nanowheel.** The experimentally measured magnetizability of the re-
3 cently synthesized tetra-cationic hexaporphyrin nanowheel (**c-P6·T6⁴⁺**) in Figure 1 is 4418
4 a.u.²³ Our calculations at the M06-2X/def2-TZVP level yield a magnetizability of 3188 a.u.
5 suggesting that DFT calculations using the M06-2X functional slightly underestimate the
6 magnetizability of the strongly antiaromatic molecules. Previous calculations on antiaro-
7 matic expanded porphyrins showed that density functionals with less than 50% Hartree-
8 Fock exchange tend to overestimate the strength of the paratropic ring current of strongly
9 antiaromatic molecules.⁴² A large positive magnetizability is obtained for **c-P6·T6⁴⁺**, be-
10 cause it sustains a strong paratropic ring current of -65 nA T^{-1} whose radius is 13 \AA . A larger
11 positive magnetizability can be obtained by either increasing the radius of the ring current
12 or by increasing the strength of the paratropic ring current. However, the diamagnetic con-
13 tribution to the magnetizability also increases with increasing size of the molecule accord-
14 ing to Pascal's rule. Closed-shell paramagnetism relies on the balance between the para-
15 magnetic contribution from χ_{zz} and the diamagnetic contributions from the χ_{xx} and χ_{yy}
16 terms.



1 **Figure 1.** The molecular structure of the tetra-cationic hexaporphyrin nanowheel
2 (**c-P6·T6⁴⁺**).²³ The paratropic ring current circulation pathway is schematically indicated
3 with the red ring with an average radius 13 Å.

4 **4.2 Cyclo[n]carbons.** The magnetizability and ring-current strengths have been also stud-
5 ied for a series of all-carbon molecules consisting of antiaromatic cyclo[n]carbons, *i.e.*, cy-
6 clic C_n polyynes (n=4k, with k=2-11) that satisfy Hückel's rule for antiaromaticity. The
7 obtained results are summarized in Table 1. The calculations show that the strength of the
8 induced paratropic ring current decreases with increasing number of carbon atoms and be-
9 comes diatropic for the two largest all-carbon rings. Since the radius of the molecular ring
10 increases with the number of carbons, the product of the radius times the ring current
11 strength passes through a maximum for n=20 (k=5), which also has the largest χ_{zz} value of
12 105 a.u. The magnetizability is positive for k=2-6, whereas the larger polyynes are diamag-
13 netic, because χ_{xx} and χ_{yy} are negative and they increase in absolute value linearly with
14 the size of the molecule. The calculations show that cyclo[4k]carbons with k=2-6 are closed-
15 shell paramagnetic molecules. Cyclo[28]carbon is diamagnetic even though it has a radius
16 of 5.7 Å and sustains a strong paratropic ring current of -15 nA T⁻¹, because the sum of the
17 diamagnetic contributions from χ_{xx} and χ_{yy} of -84 a.u. is larger than the paramagnetic χ_{zz}
18 contribution of 67 a.u. Cyclo[44]carbon molecule, which is the largest all-carbon ring that
19 we studied, is slightly aromatic with a ring-current strength of 4 nA T⁻¹ resulting in dia-
20 magnetism with a magnetizability of -59 a.u.

1 Annulenes with $(4k+2)$ π electrons exhibit a more stable π system with Hückel topology
2 and those with $(4k)$ π electrons prefer a Möbius twisted structure.⁴³ Cyclo[n]carbons ($n=4k$)
3 become aromatic for large n values due to the change from Hückel to Möbius topology of
4 the conjugated π electrons. Constructing Frost-Musulin mnemonic diagrams shows that cy-
5 clo[$4n$]carbons can be actually considered as double [n]annulenes, whose π electrons are
6 oriented orthogonally to the molecular plane and in the plane, respectively.⁴⁴⁻⁴⁶ With in-
7 creasing ring size (n), the difference between the Frost-Musulin mnemonic of the Hückel
8 system (the vertex of the n -gon is its lowest point) and the one for the Möbius system (the
9 lowest edge is horizontal) becomes smaller leading to almost the same π -electron energy.
10 For $n=36$ ($k=9$) and $n=40$ ($k=10$), the cyclo[n]carbons are non-aromatic sustaining very
11 weak ring currents of -0.5 and 0.9 nA T⁻¹, respectively, implying that the antiaromatic
12 Hückel ($4k$) and the aromatic Möbius ($4k$) structures are degenerate. For $n=44$ ($k=11$), Mö-
13 bius aromaticity clearly dominates over the Hückel antiaromaticity leading to a diatropic
14 ring current of 4.0 nA T⁻¹.



15
16 **Figure 2.** The formal molecular structures of the studied cyclo[n]carbons ($n=4k$). The
17 smallest studied molecule is cyclo[8]carbon ($k=2$) and the largest one is cyclo[44]carbon
18 ($k=11$).

1 **Table 1.** The ring-current strength (J in nA T⁻¹), the radius of the ring current (R in Å), and
2 the average magnetizability (χ in a.u.) of the studied cyclo[n]carbons.

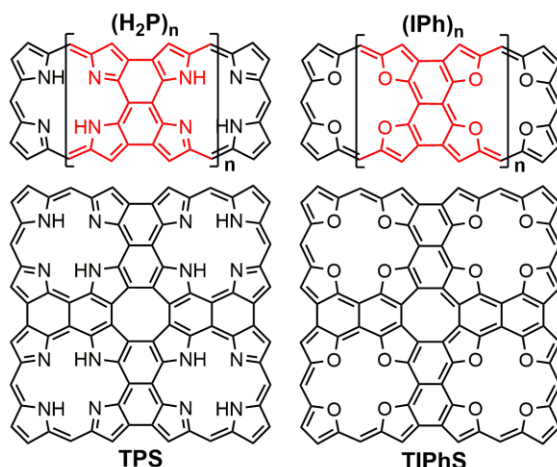
n	J	R	χ
8	-42	1.8	5
12	-40	2.4	14
16	-36	3.5	18
20	-27	4.1	16
24	-18	4.9	7
28	-15	5.7	-6
32	-4.4	6.5	-20
36	-0.5	7.4	-34
40	0.9	8.4	-47
44	4.0	9.1	-59

3 **4.3 Porphyrinoid arrays.** Free-base porphyrin (**H₂P**)₀ is an aromatic and diamagnetic mol-
4 ecule with $\chi = -68$ a.u.^{24,25,47-49} It sustains a global diatropic magnetically induced ring cur-
5 rent of 27 nA T⁻¹ by the 26 delocalized π electrons.²⁸ Isophlorin (**IPh**)₀ is chemically stable
6 and strongly antiaromatic sustaining a ring current of -27 nA T⁻¹ by its 28 π electrons.²⁴
7 Isophlorin is weakly paramagnetic with a positive magnetizability of 12 a.u. at the M06-2X
8 level of theory and of 16 a.u. at the MP2 level.²⁴ We have investigated how the magnetiz-
9 ability changes when extending porphyrin and isophlorin into the one-dimensional (1D)
10 *meso-meso*, β - β , β - β -linked arrays shown in Figure 3. In 2001, Osuka *et al.*, synthesized
11 and fully characterized a *meso-meso*, β - β , β - β -linked porphyrin consisting of twelve Zn-
12 porphyrin (**ZnP**) moieties ($n=11$)⁵⁰. Eight years later same group synthesized fully conju-
13 gated porphyrin tapes consisting of 24 **ZnP** moieties.⁵¹
14 We have studied the current density of different moieties of the 1D (**IPh**) _{n} arrays, because
15 total ring-current strengths are less meaningful. We found that the **IPh** arrays have a very

1 unusual mixed aromatic-antiaromatic character. Analysis of the current-density pathways
2 in the **IPh** arrays indicates the presence of a strong almost independent paratropic ring cur-
3 rent of -24 nA T^{-1} inside each **IPh** moiety. A weak diatropic ring current is sustained by the
4 naphthalene ring between the two **IPh** moieties. Longer **(IPh)_n** arrays ($n=2-7$) also sustain
5 localized paratropic ring currents in each **IPh** moiety. However, the paratropic ring currents
6 in the inner **IPh** fragments are somewhat weaker than the ones at the end of the **IPh** moie-
7 ties. The mixed aromatic-antiaromatic character suppresses the global electron delocaliza-
8 tion in the **(IPh)_n** arrays, which is supported by the almost size-independent HOMO-LUMO
9 gap (HLG) of the **(IPh)_n** arrays (Table 2). Thus, there is no quantum confinement effect for
10 longer **(IPh)_n** arrays, which is unexpected but can be explained in terms of alternating local
11 diatropic and paratropic ring currents. The calculations also show that the **(IPh)_n** arrays have
12 several quasi-degenerated low-lying excited states that are accessible through magnetic-
13 dipole transitions.

14 Calculations on the **(H₂P)_n** arrays show that local paratropic ring currents are confined in
15 the molecular moieties between the porphyrin subunits (Figure 4), whereas the porphyrin
16 rings sustain weak diatropic ring currents. The global delocalization of the π electrons leads
17 to a small non-vanishing diatropic ring current of 6 nA T^{-1} that flows along the central β - β
18 C-C link, which results in a smaller HLG for the longer **(H₂P)_n** arrays. The small HLG leads
19 to a stronger paratropic ring currents than for the **(IPh)_n** arrays that do not sustain any ring-
20 current flux along the central β - β C-C link, which is denoted with the dashed line in Figure
21 4. The weak global ring current of 6 nA T^{-1} for the **(H₂P)_n** arrays suggests that they have a

1 global electron delocalization leading to a decreasing HLG for longer arrays in contrast to
2 the $(\mathbf{IPh})_n$ arrays, which do not sustain any global ring current. Thus, the $(\mathbf{IPh})_n$ arrays lack
3 global electron delocalization, which results in a size-independent HLG.



4

5 **Figure 3.** The molecular structures of the studied porphyrin ($\mathbf{H}_2\mathbf{P}$) and isophlorin (\mathbf{IPh})
6 arrays. Top chart represents linear oligomers $(\mathbf{H}_2\mathbf{P})_n$ and $(\mathbf{IPh})_n$ ($n=0-7$), while in the bottom
7 part shows the square-shaped **TPS** and **TIPhS** arrays.

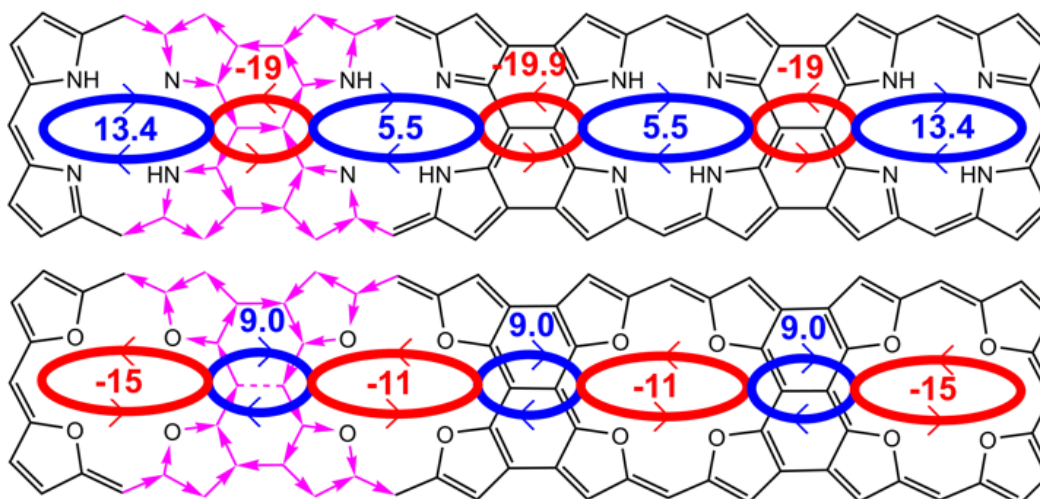
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9 The $(\mathbf{H}_2\mathbf{P})_n$ arrays are diamagnetic because the strengths of the local paratropic ring currents
10 in the naphthalene-pyrrole moieties are less than 20 nA T^{-1} , which has been found to be the
11 critical ring-current strength for closed-shell paramagnetism of porphyrinoids. The dia-
12 tropic ring currents in the porphyrin rings also cancel the paramagnetic contribution to χ_{zz} .
13 The $(\mathbf{IPh})_n$ species exhibit a diamagnetic character due to same reason namely, the strength
14 of the local paratropic ring currents in the isophlorin fragments is less than threshold value
15 of -20 nA T^{-1} and diatropic ring currents in the naphthalene moieties between the isophlorin
16 rings cancel the paramagnetic contribution to χ_{zz} . However, the magnetizability of $(\mathbf{IPh})_n$

1 is less diamagnetic than for the $(\mathbf{H}_2\mathbf{P})_n$ analogues (-11 vs. -96 a.u. for $n=1$, -127 vs. -30 a.u.
 2 for $n=2$, etc., see Table 2). Thus, if one would be able to slightly increase the strength of the
 3 local paratropic ring current in each isophlorin unit, the whole array might become para-
 4 magnetic.

5
 6 **Table 2.** The magnetizability (χ in a.u.) and the HOMO-LUMO gap (HLG in eV) of the
 7 studied porphyrin $(\mathbf{H}_2\mathbf{P})_n$ and isophlorin $(\mathbf{IPh})_n$ arrays. HLG represents the smallest denom-
 8 inator in the expression for χ_p . The ring-current strengths of porphyrin and isophlorin are 27
 9 nA T⁻¹ and -29 nA T⁻¹, respectively.

	$(\mathbf{H}_2\mathbf{P})_n$		$(\mathbf{IPh})_n$	
n	χ	HLG	χ	HLG
0	-68	4.46	12	3.39
1	-96	2.82	-11	2.22
2	-127	2.16	-30	2.17
3	-158	1.77	-48	2.14
4	-190	1.50	-68	2.14
5	-221	1.28	-87	2.12
6	-252	1.06	-106	2.12
7	-283	0.61	-124	2.11



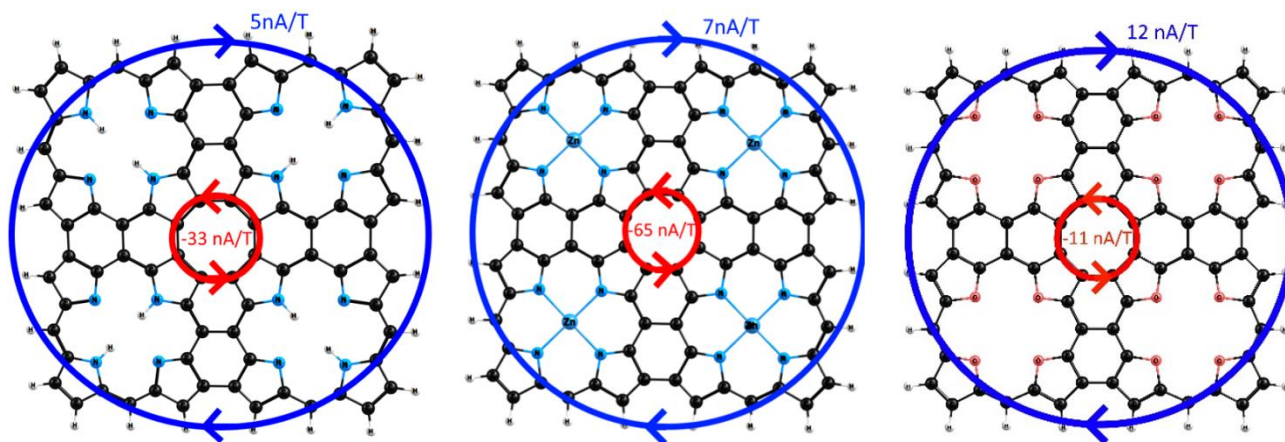
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11

1 **Figure 4.** The strong local paratropic ring currents (red ellipses) and the weak diatropic ring
2 currents (blue ellipses) in the (**H₂P**)₃ and (**IPh**)₃ arrays (top and bottom, respectively). The
3 structure of the current-density flow is the same for the whole series of the studied mole-
4 cules (n=1-7). The pink arrows show the current-density pathways, and numbers correspond
5 to strength of the ring currents (in nA T⁻¹).

6 We also studied the current density and magnetizability of two-dimensional (2D) quadratic
7 **H₂P** (**TPS**) and **IPh** (**TIPhS**) sheets (Figure 5) in which the building blocks are connected
8 by the same *meso-meso*, β - β , β - β -linking scheme as in the linear arrays. Free-base **TPS**
9 has been successfully synthesized and characterized by Osuka *et.al.*,^{52,53} while **TIPhS** is
10 predicted here. **TPS** and **TIPhS** are predicted to sustain ring currents along the outer edge
11 (rim) of the molecule and in the inner eight-membered ring (hub). In contrast to the (**H₂P**)_n
12 and (**IPh**)_n arrays, no multiple paratropic ring-current domains were found in the quadratic
13 2D sheets. The calculations show that the ring current of the hub of **TPS** and **TIPhS** is
14 paratropic (Figure 5), which agrees with previous NICS studies on **TPS**⁵³ and with studies
15 showing that hetero[8]circulenes also generally sustain a strong paratropic ring current in
16 the hub.⁵⁴⁻⁵⁶ Complexation of free-base **TPS** with Zn²⁺ yields tetra-Zn-porphyrin
17 **Zn₄TPS**^{57,58}, where the eight inner hydrogens are replaced by four Zn²⁺ cations. The com-
18 plexation enhances the paratropic ring currents in the hub increasing the ring-current
19 strength from -33 nA T⁻¹ for free-base **TPS** to -65 nA T⁻¹ for **Zn₄TPS**. The diatropic ring
20 current of 5 nA T⁻¹ and 7 nA T⁻¹ along the rim of **TPS** and **Zn₄TPS**, respectively, are much
21 weaker than the paratropic ones in the hub (Figure 5). However, due to the large radius of

1 the diatropic ring current of the rim, they play a crucial role for the sign of the magnetiza-
2 bility of **TPS** and **Zn₄TPS**. The paramagnetic contribution from the strong paratropic ring
3 current of the hub with a small radius is cancelled by the diamagnetic contribution from the
4 weak diatropic ring current with a large radius. In addition, each atom brings diamagnetic
5 contributions to the magnetizability according to Pascal's rule. Therefore, free-base **TPS**
6 and the corresponding **Zn₄TPS** complex are diamagnetic with magnetizabilities of -126 a.u.
7 and -132 a.u., respectively. Due to the same reason, **TIPhS** sustains a paratropic ring current
8 of -11 nA T⁻¹ in the hub and a diatropic ring current of 12 nA T⁻¹ along the rim leading to a
9 diamagnetic molecule with a negative magnetizability of -140 a.u.



11 **Figure 5.** The diatropic (blue line) and paratropic (red line) magnetically induced ring cur-
12 rents in the free-base tetra-porphyrin **TPS** sheet (left), the **Zn₄TPS** sheet (middle), and in
13 the tetra-isophlorin **TIPhS** sheet (right).

14 **4.4 Iso[n]phlorins.** The magnetizability and ring-current strengths were calculated for iso-
15 phlorin **IPh** (iso[4]phlorin) and a series of analogous expanded isophlorins (Figure 6). The
16 molecular charge has been chosen such that the expanded isophlorins fulfill Hückel's rule

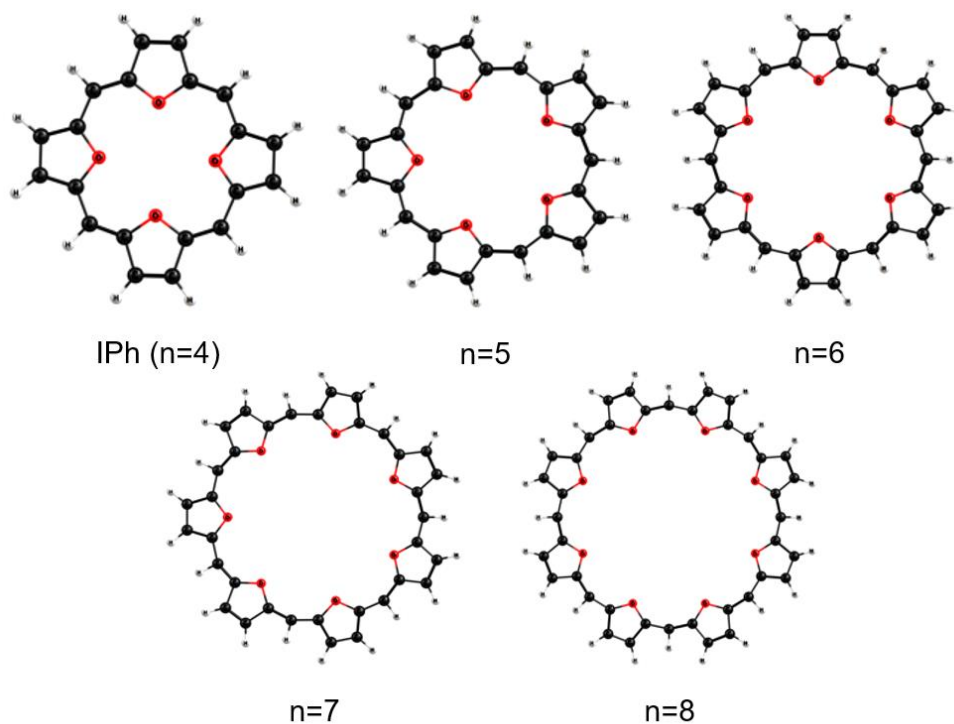
1 for antiaromaticity. Isophlorin is an antiaromatic porphyrin with 28 π electrons in the por-
2 phyrin ring. Originally isophlorin was constructed from free-base porphyrin by adding two
3 inner hydrogens.⁵⁹ Alternatively, one can replace the NH moieties with iso-electronic ox-
4 ygen atoms.⁶⁰⁻⁶² The expanded isophlorins studied here consisting of 5-8 furan rings have
5 not been synthesized, whereas the synthesized cyclo-octafuran(1.1.1.1.1.1.1.1) has in-
6 verted furan rings.⁶³ Our calculations show that the inverted (experimental) configuration
7 is 210 kcal mol⁻¹ more favourable than a hypothetical circle-shaped iso[8]phlorin. For the
8 design of closed-shell paramagnetic antiaromatic molecules, we have considered a series
9 of iso[n]phlorins of the same circle-shaped topology. Calculations on this series of mole-
10 cules yield a relation between the ring-current strength, the ring-current radius and the
11 magnetizability. We have previously showed that the isophlorin ring consisting of four
12 furan rings connected via methine groups is strongly antiaromatic sustaining a paratropic
13 ring current of -63 nA T⁻¹ at the B3LYP level.⁴⁸ The present M06-2X calculations yield a
14 ring-current strength of -29 nA T⁻¹ for isophlorin. The magnetizability of isophlorin calcu-
15 lated at the M06-2X level is 12 a.u. The calculated ring-current strengths and magnetiza-
16 bilities of the antiaromatic ($n=4k$) expanded isophlorins are summarized in Table 3. The
17 studied isophlorins sustain strong paratropic ring currents that are of about the same
18 strength except for iso[8]phlorin, which sustains a very strong paratropic ring current of -
19 617 nA T⁻¹. Since the radius increases with the number of furan rings, the magnetizability
20 also increases with n . For iso[8]phlorin the large radius of the ring current of 5.3 Å and the
21 strong paratropic ring current lead to a very large positive magnetizability of 3532 a.u.

22

1 **Table 3.** The ring-current strength (J in nA T⁻¹), the radius of the ring current (R in Å), and
 2 the average magnetizability (χ in a.u.) of the studied iso[n]phlorins, where n is the number
 3 of furan rings.

Molecule	J	R	χ
Iso[4]phlorin	-29	2.0	12
Iso[5]phlorin ¹⁺	-73	2.5	132
Iso[6]phlorin ²⁺	-56	3.7	140
Iso[7]phlorin ¹⁻	-45	4.5	158
Iso[8]phlorin	-617	5.3	3532 ^a

4 ^a For comparison, the magnetizability of **c-P6·T6**⁴⁺ is 4418 a.u.²² and the magnetizability of O₂ (³ Σ_g) is
 5 719 a.u. at a temperature of 293 K.⁶⁴



6
 7 **Figure 6.** The molecular structures of the studied iso[n]phlorins, $n=4-8$.

8 **4.5. Expanded porphyrinoids.** We also studied orangarin, rosarin, and amethyrin, which
 9 are synthesized expanded antiaromatic porphyrins whose molecular structures are shown
 10 in Figure 7.⁶⁵ Even though they are strongly antiaromatic,⁴¹ calculations of the magnetiza-
 11 bilities show that they are diamagnetic (Table 4). The paratropic ring currents are weaker

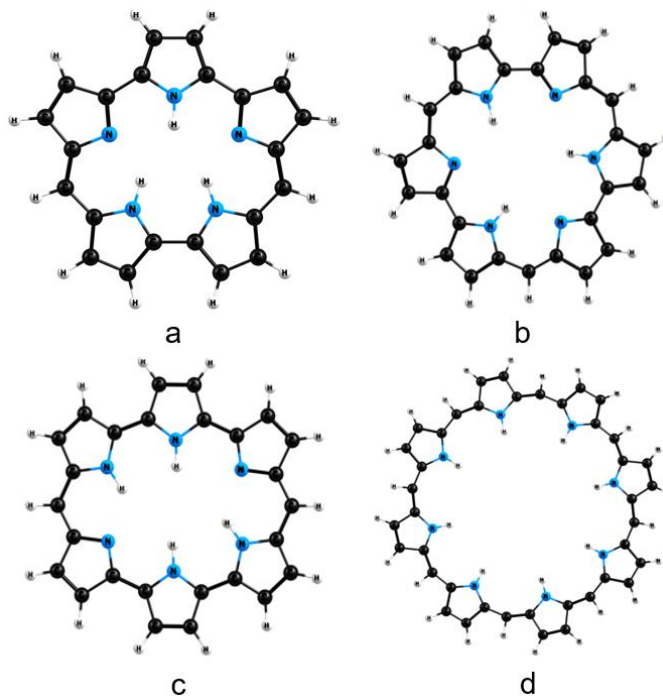
1 than -20 nA T^{-1} and the radius of the ring current is not large enough to lead to closed-shell
2 paramagnetism.

3 We also studied the hypothetical circle-shaped [40]octaphyrin(1.1.1.1.1.1.1.1) molecule
4 with eight NH groups that consists of eight pyrrole rings connected *via* methine bridges.
5 The molecular structure of [40]octaphyrin(1.1.1.1.1.1.1.1) shown in Figure 7d is a local
6 minimum on the potential energy surface of the singlet ground state; it lies $163 \text{ kcal mol}^{-1}$
7 above the square-shaped [40]octaphyrin(1.1.1.1.1.1.1.1) shown in Figure 8b also contain-
8 ing eight NH groups, which has not been synthesized. [40]octaphyrin(1.1.1.1.1.1.1.1) with
9 inverted pyrrole rings is diamagnetic with a negative magnetizability of -67 a.u. and prac-
10 tically nonaromatic sustaining a ring current of $J=3.4 \text{ nA T}^{-1}$. In contrast, the circle-shaped
11 [40]octaphyrin(1.1.1.1.1.1.1.1) sustains a strong paratropic ring current of -561 nA T^{-1} re-
12 sulting in a large positive magnetizability of 3244 a.u.

13 **Table 4.** The ring-current strength (J in nA T^{-1}), the radius of the ring current (R in \AA), and
14 the average magnetizability (χ in a.u.) of orangarin, rosarin, amethyrin and the hypothetical
15 [40]octaphyrin(1.1.1.1.1.1.1.1).

Molecule	J	R	χ
Orangarin	-20	2.0	-7.0
Rosarin	-18	2.5	-4.9
Amethyrin	-14	2.5	-14
[40]octaphyrin(1.1.1.1.1.1.1.1) (circle-shaped)	-561	5.3	3244

17



1

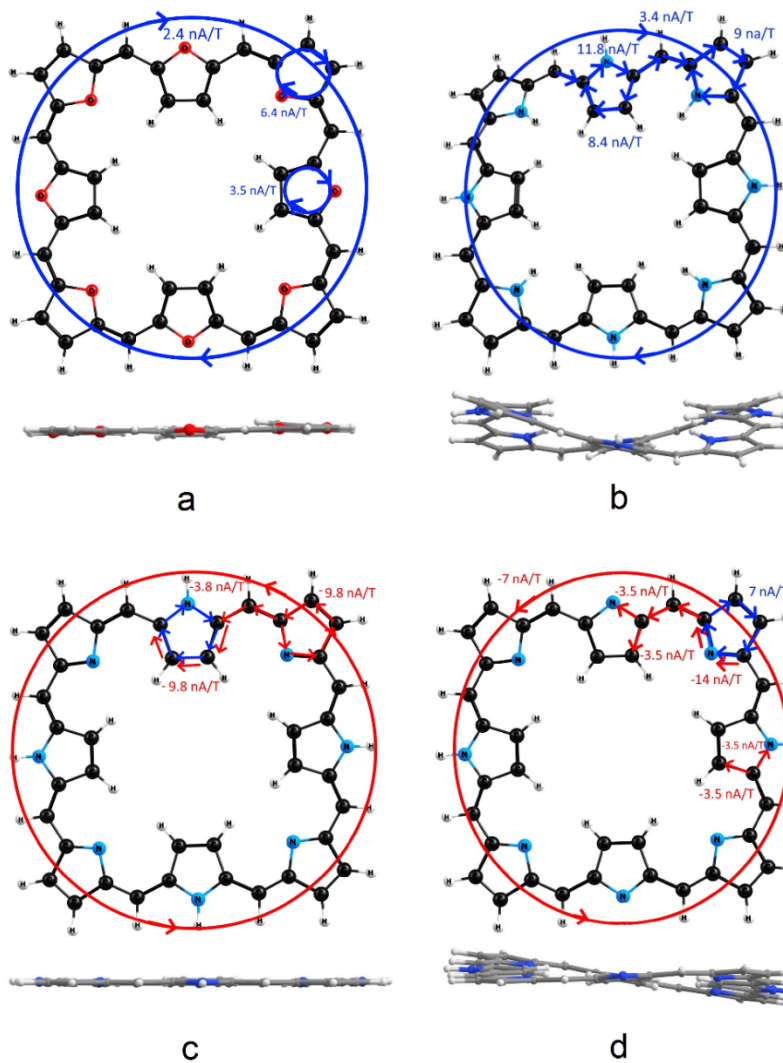
2 **Figure 7.** The molecular structure of the antiaromatic expanded porphyrins: a) orangarin
 3 b) rosarin c) amethyrin and d) circle-shaped [40]octaphyrin(1.1.1.1.1.1.1.1).

4

5 The anomalously high paratropic ring currents for the hypothetical iso[8]phlorin in Figure
 6 6 and [40]octaphyrin(1.1.1.1.1.1.1.1) in Figure 7d seem to originate from the circular to-
 7 pology of the molecular structure. They are examples of closed-shell paramagnetic mole-
 8 cules with a strong ring current and a large ring-current radius. Since iso[8]phlorin and
 9 circle-shaped [40]octaphyrin(1.1.1.1.1.1.1.1) are strongly antiaromatic, they are high-lying
 10 local minima on the potential energy surfaces of the singlet ground state. The global min-
 11 imum structure has four pyrrole/furan rings with the NH/O moiety pointing inwards and
 12 four NH/O moieties pointing outwards. The synthesis of the high-lying structures is cer-
 13 tainly challenging, whereas from the electronic structure point of view these molecules are
 14 stable.

1 **4.6. Square-shaped [36]octaphyrin(1.1.1.1.1.1.1) and cyclo-octafuran(1.1.1.1.1.1.1)**

2 The previously synthesized [36]octaphyrin(1.1.1.1.1.1.1)⁶⁶ and cyclo-octafu-
3 ran(1.1.1.1.1.1.1)⁶³ with inverted rings shown in Figure 8 are weakly aromatic and
4 weakly antiaromatic molecules sustaining a weak diatropic (2.4 nA T⁻¹) and weak para-
5 tropic (-3.8 nA T⁻¹) global ring current, respectively. Cyclo-octafuran (1.1.1.1.1.1.1)⁶³ is
6 a planar diamagnetic molecule with a magnetizability of -56 a.u. It sustains local diatropic
7 ring currents of 5.9 and 8.8 nA T⁻¹ in the furan rings. The [36]octaphyrin(1.1.1.1.1.1.1)⁶⁶
8 molecule shown in Figure 8c (it contains four NH groups in contrast to hypothetical
9 [36]octaphyrin(1.1.1.1.1.1.1) in Figure 8b) is also planar and diamagnetic molecule with
10 a magnetizability of -9.7 a.u. It is a weakly antiaromatic 40 π electron annulene, since the
11 N/NH moieties do not contribute to the global ring current due to the local ring currents in
12 the pyrrole rings. When two protons are removed from [36]octaphyrin(1.1.1.1.1.1.1)
13 (Figure 8d), the resulting dianion is antiaromatic with a ring-current strength of -7 nAT⁻¹
14 and paramagnetic with a magnetizability of $\chi=9$ a.u. It is a porphyrinoid with 46 π electrons
15 showing that large molecular rings sustaining relatively weak paratropic ring currents can
16 be closed-shell paramagnetic molecules. The optimized molecular structure of the [36]oc-
17 taphyrin(1.1.1.1.1.1.1) dianion is ruffled, while in the single-crystal X-ray diffraction
18 measurement, the dianion with C₆F₅ substituents in the *meso*-positions and with tetrabu-
19 tylammonium counterions is nearly planar. Thus, we propose that the unsubstituted [36]oc-
20 taphyrin(1.1.1.1.1.1.1) dianion is a Möbius-antiaromatic 46 π electron molecule.



1

2 Figure 8. Current-density pathways and ring-current strengths calculated for (a) cyclo-oc-
 3 tafuran(1.1.1.1.1.1.1.1), (b) hypothetical square-shaped [40]octaphyrin(1.1.1.1.1.1.1.1),
 4 (c) [36]octaphyrin(1.1.1.1.1.1.1.1) and (d) [36]octaphyrin(1.1.1.1.1.1.1.1) dianion using
 5 GIMIC. The side view of the molecular structures are also shown.

6 CONCLUSIONS

7 The present study gives a recipe for designing closed-shell paramagnetic molecular rings.
 8 One of the prerequisites for closed-shell paramagnetism (CSP) of ring-shaped porphyrinoids
 9 seems to be that they sustain a paratropic ring current of at least -20 nA T^{-1} . The paratropic

1 ring current must not be cancelled by other diatropic ring currents with larger radii. Larger
2 radii of the paratropic ring current lead to stronger closed-shell paramagnetism. However,
3 the diamagnetic contribution to the magnetizability also increases with the size of the ring
4 according to Pascal's rule. For planar molecules in the xy plane, the zz component of the
5 magnetizability tensor can be positive due to a strong paratropic magnetically induced ring
6 current, whereas the xx and yy components are negative (diamagnetic) according to Pascal's
7 rule, since planar molecules do not sustain any paratropic ring currents when the magnetic
8 field is parallel to the molecular plane. Thus, closed-shell paramagnetism is a balance be-
9 tween the paramagnetic contribution to the zz element of the magnetizability tensor and the
10 diamagnetic contributions from the xx and yy elements.

11 The reliability of the employed computational level was assessed by calculating the magnet-
12 ization of the recently synthesized tetra-cationic hexaporphyrin nanowheel (**c-P6·T6⁴⁺**),
13 which is a closed-shell paramagnetic molecule with an experimental magnetizability of 4418
14 a.u.²³ Calculations at the M06-2X yielded a magnetizability of 3188 a.u. showing that the
15 employed computational level is able to reproduce the large positive magnetizability of **c-**
16 **P6·T6⁴⁺**. Magnetizabilities calculated at the M06-2X level for closed-shell paramagnetic
17 molecular rings seem though to be slightly underestimated.

18 Iso[4]phlorin sustaining a ring current of -29 nA T^{-1} with a radius of 2 \AA is paramagnetic
19 with a magnetizability of 12 a.u. Cyclo[8]carbon with a radius of 1.8 \AA sustaining a ring
20 current of -42 nA T^{-1} is weakly paramagnetic with a magnetizability of 5 a.u. These examples
21 show that closed-shell paramagnetic molecular rings with a radius of $< 2 \text{ \AA}$ must sustain a
22 very strong paratropic ring current.

1 A series of *meso-meso*, β - β , β - β -linked porphyrins (**TPS**) and isophlorins (**IPh**) extending
2 in one (arrays) and two (sheets) dimensions were studied. The linear and square-shaped
3 porphyrin and isophlorin sheets have a different topology of the ring currents. The linear
4 porphyrinoid arrays have multiple centers sustaining local diatropic and paratropic ring cur-
5 rents, whereas the squared-shaped porphyrinoid sheets sustain a paratropic ring current in
6 the hub and a diatropic ring current along the rim. Substituting the inner NH protons of **TPS**
7 with Zn^{2+} cations leads to a stronger paratropic ring current in the hub, whereas the diatropic
8 ring current along the rim is less affected.

9 The linear isophlorin arrays (**IPh**)_n sustain paratropic ring currents in each of the isophlorin
10 moieties and a diatropic ring current in the naphthalene moieties between the isophlorins.
11 The (**IPh**)_n arrays do not sustain any global ring current. The absence of global electron
12 delocalization leads to a very weak quantum confinement for the (**IPh**)_n arrays implying that
13 the HOMO-LUMO gap is almost independent of the number of isophlorins moieties for the
14 longer (**IPh**)_n arrays.

15 For the (**H₂P**)_n arrays, a global diatropic ring current through the β - β links results in a smaller
16 HOMO-LUMO gap for the longer (**H₂P**)_n arrays. The porphyrin moieties sustain local dia-
17 tropic ring currents, whereas the linking naphthalene-pyrrole moieties sustain local para-
18 tropic ring currents. The linear arrays sustaining several local paratropic ring currents are
19 diamagnetic, because the ring-current strengths do not exceed -20 nA T^{-1} . Several paratropic
20 ring currents in the same molecule does not make them paramagnetic, because the diamag-
21 netic contribution increases also linearly with the number of atoms in the molecule.

1 Calculations on a series of expanded isophlorins fulfilling Hückel's rule for antiaromaticity
2 show that they are paramagnetic due to the strong paratropic ring current and the large radius
3 of the ring current. The largest member of the series namely iso[8]phlorin has a magnetiza-
4 bility of 3532 a.u., because it sustains a very strong ring current of -617 nA T^{-1} whose radius
5 is 5.3 \AA . Calculations on expanded porphyrins fulfilling Hückel's rule for antiaromaticity
6 show that orangarin, rosarin, amethyrin are diamagnetic, since they sustain paratropic ring
7 currents that are weaker than -20 nA T^{-1} . The largest member of that series is the hypothetical
8 circle-shaped [40]octaphyrin(1.1.1.1.1.1.1.1) that sustains a strong paratropic ring current of
9 -561 nA T^{-1} whose radius is 5.3 \AA . The circle-shaped [40]octaphyrin(1.1.1.1.1.1.1.1) mole-
10 cule is therefore strongly paramagnetic with a magnetizability of 3244 a.u.

11

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15 **Notes**

16 The authors declare no competing financial interests.

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7

8 **ASSOCIATED CONTENT**

9 **Supporting Information**

10 The Supporting Information is available free of charge on the ACS Publications website at
11 DOI: 10.1021/acs.jpcc.XXXXXXX. Optimized geometries (Cartesian coordinates) for all
12 studied molecules (c-P6·T6⁴⁺ nanowheel, cyclo[n]carbons (n=8, 12, 16 ... 44), (H₂P)_n (n=0-
13 7) and (IPh)_n (n=0-7) arrays, TPS, Zn₄TPS and TPhS nanosheets, iso[n]phlorins (n=4-8),
14 expanded porphyrinoids, octaphyrins and cyclo-octa furans).

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